JOURNAL OF CATALYSIS 141, 321–322 (1993)

ERRATUM

Volume **135,** Number 2 (1992), in the article "Piperidine Hydrogenolysis on a Commercial Hydrocracking Catalyst. III. The Effects of Zeolite Unit Cell Size, Catalyst Sulfur Content, and Coke Deposition on Catalyst Activity and Deactivation," by G. C. Hadjiloizou, J. B. Butt, and J. S. Dranoff, pages 481–504: Equation (5) of the paper should read

$$r_i = k_i^{\circ} C_1^2 s_i = \frac{k_i^{\circ\prime\prime} K_{\mathrm{pip}_i}{}^2 C_1^2 s_i}{(1 + K_{\mathrm{pip}_i} C_{1_0})^2}. \qquad (5)$$

Reference (*10*) should read

10. Hadjiloizou, G. C., Butt, J. B., and Dranoff, J. S., *J. Catal.* **135,** 27 (1992).

The following Nomenclature section was inadvertently omitted.

NOMENCLATURE

a_i	Decay parameter for the formation of product i, $\mathrm{min}^{-0.5}$.
C_1	Concentration of piperidine, g mol/liter.
C_{1_0}	Initial concentration of piperidine, g mol/liter.
D	Average deviation, defined as $D = (1/m)\ \Sigma_m\ \{[\lvert k''_{i,\mathrm{obs},m} - k''_{i,\mathrm{pred},m}\rvert]/k''_{i,\mathrm{obs},m}\} \times 100$.
E_{d_t}	Activation energy of overall deactivation, kcal/g mol.
F_{1_0}	Reactor inlet molar flow rate of piperidine, g mol/min.
i	Index: $i = 2$ = *N*-*n*-pentylpiperidine, $i = 3$ = 2-*n*-pentylpiperidine, $i = 4$ = decahydroquinolines.
K_{pip_i}	Adsorption equilibrium constant for piperidine associated with formation of product i, liter/g mol.
K'_{pip_i}	Preexponential factor of K_{pip_i}, liter/g mol.
k_{d_t}	Rate constant of overall deactivation reaction, liter/g mol/min.
k'_{d_t}	Preexponential factor of k_{d_t}, liter/g mol/min.
$k^*_{d_t}$	Apparent rate constant of overall deactivation reaction, min^{-1}.
k_i	Reaction rate constant of formation of product i.
k_i°, $k_i^{\circ\prime\prime}$	Initial reaction rate constants of formation of product i, liter^2/g mol/g cat/min and g mol/g cat/min, respectively.
$k^{*\prime}_{ij}$, $k^{\circ\prime\prime}_{ij}$	Rate constants of formation of product i on catalyst j, as defined in Eq. (6).
k_t°	Initial rate constant of overall reaction, liter^2/g mol/g cat/min.
N_{ij}°	Initial total number of sites active for formation of product i on catalyst j.
n^*	Order of concentration dependency of deactivation.
pip	Piperidine.
R	Gas constant, kcal/g mol/K.
R^2	Coefficient of determination for regression analysis.
r_i	Reaction rate of formation of product i, g mol/g cat/min.
s_i	Activity of catalyst at time on stream t for formation of product i.
T	Absolute temperature, K.

t Time on stream, min.

W Weight of catalyst, g.

x_i Conversion of piperidine to product i at time on stream t.

x_i^o Initial conversion of piperidine to product i.

x_t Total conversion of piperidine at time on stream t.

x_t^o Initial total conversion of piperidine.

Greek Symbols

ΔH_{pip_i} Enthalpy change of adsorption of piperidine associated with formation of product i, kcal/g mol.

τ Space time $= WC_{\text{I}_o}^2/F_{\text{I}_o}$, g mol · g cat · min/liter^2.